

Optimization and evaluation of CBSOL LE red wool dye adsorption from aqueous solution onto commercial activated carbon

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Abstract Adsorptive removal of CBSOL LE red wool dye from aqueous solution onto commercial activated carbon (CAC) was investigated in a batch system. Various process parameters like pH, dosage of CAC (m) and adsorption time (t) were considered and optimized with full factorial central composite design under response surface methodology. At optimized parameters, kinetic and thermodynamic studies were performed and adsorption equilibrium data were represented using Freundlich, Langmuir and Redlich–Peterson (R–P) isotherm models. Also, diffusivity was calculated for the rate-limiting step in the adsorption process. Optimum process parameters were found to be $m = 2.92$ g/100 ml, $t = 6.75$ h and $\text{pH} = 3.95$, and at these optimized parameters, % removal of CBSOL LE red wool dye (Y) was found to be 86 %. Pseudo-second-order kinetic was found to best fit the adsorption kinetic data. Freundlich and R–P isotherms were found to best represent the equilibrium adsorption data. Diffusivity for the intra-particle diffusion was found to be 9.676×10^{-8} and 1.396×10^{-8} m²/s at initial concentration of CBSOL LE red wool dye (C_0) of 50 and 100 mg/l, respectively.

Keywords CBSOL LE red wool dye · Commercial activated carbon · Adsorption kinetics · Isotherms · Diffusivity · Response surface methodology

Introduction

Colored effluents produced from the textile/dyeing industries are discharged into water bodies since the dyeing technique was invented (Chiou and Chuang 2006). More than 100,000 commercially available dyes with over 7×10^5 tonnes of dyestuff are produced and used annually (Ramaraju et al. 2013). Most of these dyes belong to synthetic dyes that present increased environmental hazard due to being toxic in nature (Lemlikchi et al. 2012).

CBSOL LE dyes are reactive dyes developed for wool dyeing (<http://www.colorbandindia.com/cbsol-le-reactive-dyes-for-wool.html>). Reactive dyes are azo-based chromophores combined with different types of reactive groups and are most commonly used type of dyes. Also, reactive dyes are non-biodegradable and can be carcinogenic to human (Gregory 1986). During the textile processing, large amount of dyes stay behind in wastewater produced, due to which textile wastewater becomes strong in color and toxic (Yazdani et al. 2012). Hence, the removal of dyes from wastewater is very important from both points of view of environment as well as human health.

Due to non-biodegradability of reactive dyes, physico-chemical methods such as chemical coagulation, adsorption and membrane filtration are generally used for the treatment of dye-containing wastewater (Georgiou et al. 2002; Izares et al. 2006). Drawback of chemical coagulation is that it produces secondary pollutants, and membrane processes are very expensive, whereas adsorption process presents simple design, low-cost, easy operation and insensitivity to toxic substances (Chakraborty et al. 2012; Chowdhury and Saha 2010).

Various authors have reported reactive dye adsorption from wastewater on varieties of low-cost adsorbents like tannery sludge developed activated carbon (Geethakarthis

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and Phanikumar 2011); sugarcane bagasse ash (Gaikwad et al. 2010); groundnut shell prepared activated carbon (Malik et al. 2006); mangrove bark (Seey et al. 2012); corn cob and sunflower seed shells (Suteu et al. 2011a, b); pineapple leaf powder (Weng et al. 2009); and agricultural waste-based activated carbon (Manoochchri et al. 2010; Parvathi and Maruthavanan 2010). A good review on methylene blue adsorption by low-cost adsorbent is available in the literature (Rafatullaha et al. 2010). Among various low-cost adsorbent, commercial activated carbon (CAC) has been reported having highest adsorption capacity (Rafatullaha et al. 2010; Singh and Kushwaha 2014). Among the studies available in the literature, none of the works reports about removal of CBSOL LE red wool dye.

In present work, adsorption of CBSOL LE red wool dye on the CAC was investigated. Various adsorption process parameters like pH, dosage of CAC (m) and adsorption time (t) were considered and optimized with the help of full factorial central composite design (CCD) under response surface methodology (RSM). At optimized operational parameters, kinetics and thermodynamics of adsorption process have been studied, and change in Gibbs energy (ΔG_0), entropy (ΔS_0) and enthalpy (ΔH_0) was determined at various temperatures. To represent the adsorption equilibrium data at optimized parameters, Freundlich, Langmuir and Redlich–Peterson (R–P) isotherm models were used. Also, diffusivity was calculated for the rate-limiting step in the adsorption process.

Materials and methods

Adsorbent and adsorbate

CAC was purchased from Pneumatic Engineers Spares and Service, New Delhi, India, and it was used as procured. Detailed physicochemical characteristics of the CAC have already been presented elsewhere (Singh and Kushwaha 2014). The BET surface area and BET average pore diameter of CAC were found to be 222.33 m²/g and 26.56 Å, respectively. The average particle diameter and bulk density of CAC were found to be 1.58 mm and 540.22 kg/m³, respectively.

CBSOL red LE wool dye (supplied by Grasim industry, Nagda, Madhya Pradesh, India) was dissolved in distilled water to make the dye solution. For this, 1 g/l of stock dye solution was prepared, which was further diluted to make the working solution of desired concentration as per the experimental requirement.

Experimental design, procedure and data analysis

In this study, experiments for the adsorption of CBSOL LE red wool dye from aqueous solution on CAC were designed and conducted in accordance with CCD of Design-Expert® 6.0 under RSM. The RSM is a statistical tool which helps to design the experiments and optimize the effective parameters for desirable responses providing process models. With RSM, analysis of the interaction between parameters on responses is possible with reduced number of experiments (Soo et al. 2004). Three adsorption process parameters (pH: 3–9; m : 0.5–5 g/100 ml and t : 1–7 h) with five levels were used to design experiments using RSM, and percentage CBSOL red LE wool dye removal (Y) has been taken as a response of the system. Experimental design matrix is given in Table 1.

A total of 20 experiments were conducted as suggested by RSM. The experiments were conducted in batch mode to study the effects of the three process parameters on response Y . For each set of suggested adsorption experiments, 100 ml solution of previously adjusted pH of CBSOL red LE wool dye having 100 mg/l concentration (C_0) was taken in 250-ml stoppered conical flasks. Thereafter, CAC was added to the flask as per design matrix. Initial pH of the solutions was adjusted by using 0.1 N H₂SO₄ and 0.1 N NaOH solutions. The flasks were then kept into temperature-controlled orbital shaking incubator (Remi Elektronik Ltd., India; CIS-24 Plus), and the mixture was agitated at constant speed of 150 rpm at temperature (T) 303 K for desired time (as per design matrix) to attain equilibrium.

The adsorbent and adsorbate were separated from the mixture after desired time (as per design matrix) by filtering the mixture with Whatman no.1 filter paper, and filtrate was analyzed for residual CBSOL LE red wool dye concentration (C_e), using a double beam UV–visible spectrophotometer (Electronic Corporation India Ltd., UV 5704SS) at $\lambda_{\max} = 504$ nm. The % color removal (Y) was then calculated using the following relation:

$$Y = \frac{(C_0 - C_e) 100}{C_0} \quad (1)$$

Similarly, for the equilibrium adsorption uptake, q_e (mg/g) was calculated as:

$$q_e = \frac{(C_0 - C_e) V}{m} \quad (2)$$

where C_e is the equilibrium adsorbate concentration, V is the volume of the adsorbate (l), and m is the mass of adsorbent (g).



Table 1 Full factorial design used and responses for the % removal CBSOL red LE wool dye by CAC

pH	<i>T</i> (h)	<i>m</i> (g/100 ml)	% Dye removal (<i>Y</i>)	
			Actual	Predicted
4.5	2.5	1.625	26.94	42.13
7.5	2.5	1.625	35.36	36.17
4.5	5.5	1.625	63.07	62.10
7.5	5.5	1.625	43.33	50.52
4.5	2.5	3.875	93.71	101.22
7.5	2.5	3.875	45.35	61.02
4.5	5.5	3.875	91.78	105.67
7.5	5.5	3.875	60.34	59.85
3	4	2.75	88.23	77.76
9	4	2.75	30.23	25.99
6	1	2.75	76.94	64.70
6	7	2.75	85.97	83.50
6	4	0.5	22.30	18.54
6	4	5	97.90	86.96
6	4	2.75	85.00	80.61
6	4	2.75	84.36	80.61
6	4	2.75	82.74	80.61
6	4	2.75	81.13	80.61
6	4	2.75	81.78	80.61
6	4	2.75	83.39	80.61

The response of the system, *Y*, obtained was processed for Eq. (3) with Design-Expert® 6.0. through ANOVA to obtain the interaction between the process variables and the response.

$$Y = c_o + \sum_{i=1}^4 c_i X_i + \sum_{i=1}^4 c_{ii} X_i^2 + \sum_{i=j}^3 \sum_{i=j+1}^4 c_{ij} X_{ij} \quad (3)$$

where *Y* is response, c_o , c_i , c_{ii} , c_{ij} are constant coefficients, and X_i the independent process variables.

The fit quality of second-order polynomial model was expressed by the coefficient of determination R^2 and adjusted R^2 . For optimization, desirability function module of Design Expert 6.0 was utilized, in which individual goals are combined into an overall desirability function for maximization to find the best maximum of response (Güven et al. 2008).

Results and discussions

Model equation and validation

The response of the system, *Y*, (Table 1) obtained was processed with Design-Expert® 6.0. Sequential model sum of squares, model summary statistics and ANOVA were

Table 2 ANOVA for response surface quadratic model

Source	Sum of squares	df	Mean square	<i>F</i> value	Prob > <i>F</i>
Model	10,533.12	9	1,170.35	9.41	0.0008
A	2,680.75	1	2,680.75	21.55	0.0009
B	353.59	1	353.59	2.84	0.1227
C	4,681.84	1	4,681.84	37.64	0.0001
A ²	1,297.88	1	1,297.88	10.44	0.0090
B ²	66.67	1	66.66	0.54	0.4809
C ²	1,220.14	1	1,220.14	9.81	0.0107
AB	15.77	1	15.77	0.13	0.7292
AC	586.19	1	586.19	4.71	0.0551
BC	120.49	1	120.49	0.97	0.3482
Residual	1243.70	10	124.37		
Lack of fit	1,232.68	5	246.54	111.80	<0.0001
Pure error	11.025	5	2.20		
Cor total	11,776.82	19			

A: pH

B: *t*

C: *m*

analyzed, and it was suggested by sequential model sum of squares that quadratic model most suitably described the dye removal from aqueous solution. Suggested quadratic equation is given below:

$$Y = 80.61 - 12.94 \text{ pH} + 4.7 t + 17.11 m - 7.18 \text{ pH}^2 - 1.63 t^2 - 6.97 m^2 - 1.4 (\text{pH} \times t) - 8.56 (\text{pH} \times m) - 3.88 (t \times m) \quad (4)$$

ANOVA result is shown in Table 2. Model *F* value of 9.41 and model Prob > *F* value of 0.0008 for % dye removal imply that quadratic model fitted is significant. Value of “Prob > *F*” <0.05 indicates model terms are significant. ANOVA from the response surface models shows that pH, *m*, pH² and *m*² are significant terms for response *Y*, implying that quadratic model (Eq. 4) can be used to predict response *Y* for CBSOL red LE wool dye from the aqueous solution.

Normal % probability plot of residuals and plot of studentized residuals versus predicted were evaluated (not shown here). The normality assumption was relatively satisfied as the points in the plot form fairly straight line for response *Y*. For a reliable model, no series of increasing or decreasing points, patterns and a predominance of positive or negative residuals should be found (Behbahani et al. 2011; Bansal et al. 2013). Generally, plot of studentized residuals versus predicted should be a random scatter, which suggests that the variance of original observations is constant for all values of the response, *Y* (Özer et al. 2009).



R^2 and adjusted R^2 for a regression model are evaluated to show the model fitting to experimental data. Actual versus predicted Y (% CBSOL red LE dye removal) plot (not shown here) (actual values were calculated experimentally, while predicted Y values were evaluated using Design Expert 6) revealed the values of R^2 and adjusted R^2 to be 0.894 and 0.80, which concludes a very good correlation between the experimental and the predicted values of response Y .

The outlier t plot (not shown here) is used to measure how many standard deviations the actual value deviates from the predicted value. For a good correlation, outlier t values should fall within ± 3.50 (Özer et al. 2009; Montgomery 1996; Myers and Montgomery 2002). In present study, the outlier t values were found within the interval of ± 3.50 , concluding that the approximation of the fitted model to the response surface was fairly good.

Effect of solution pH, adsorption time (t) and dosage (m)

Values of response Y for the adsorptive removal of CBSOL red LE wool dye with various t (h), pH and m (g) values are presented in Table 1, and the corresponding response surface 3D plot is shown in Fig. 1.

It can be observed from Fig. 1a that Y value increases with decreases in pH, and maximum Y value is found at $\text{pH} \approx 5$. For all $\text{pH} > 5$, Y value changes marginally. This trend of variation in Y with pH was observed at all t values. It can also be observed from Fig. 1a that when pH of reactive CBSOL red LE wool dye aqueous solution is less than ≈ 5 , adsorption of reactive CBSOL red LE wool dye increases with increase in t ($t \leq 5$ h), and for all $t > 5$ h values, Y becomes constant. But, for $\text{pH} > 5$, increase in pH showed very minute dye adsorption at all t values.

To understand the mechanism of adsorption, study for isoelectric point (point of zero charge, pH_{PZC}) determination of adsorbent was conducted and pH_{PZC} was found to be 8.6. At pH below $\text{pH}_{\text{PZC}} = 8.6$, due to protonation, surface of adsorbent becomes positively charged (Lakshmi et al. 2009), which favors dye adsorption, and hence, Y value increases with decrease in pH, whereas at $\text{pH} > \text{pH}_{\text{PZC}}$, adsorbent surface becomes negatively charged due to ionization and dye adsorption is not favoured.

Response surface 3D plot of Y versus t and m is shown in Fig. 1b. For any t value, increasing m up to $m \approx 2.9$ g resulted in rapid increase in response Y . Further increasing m values ($m > 2.9$ g), response Y was found to gradually decrease, and being constant to $Y = 85\%$ at all $t < 5$ h

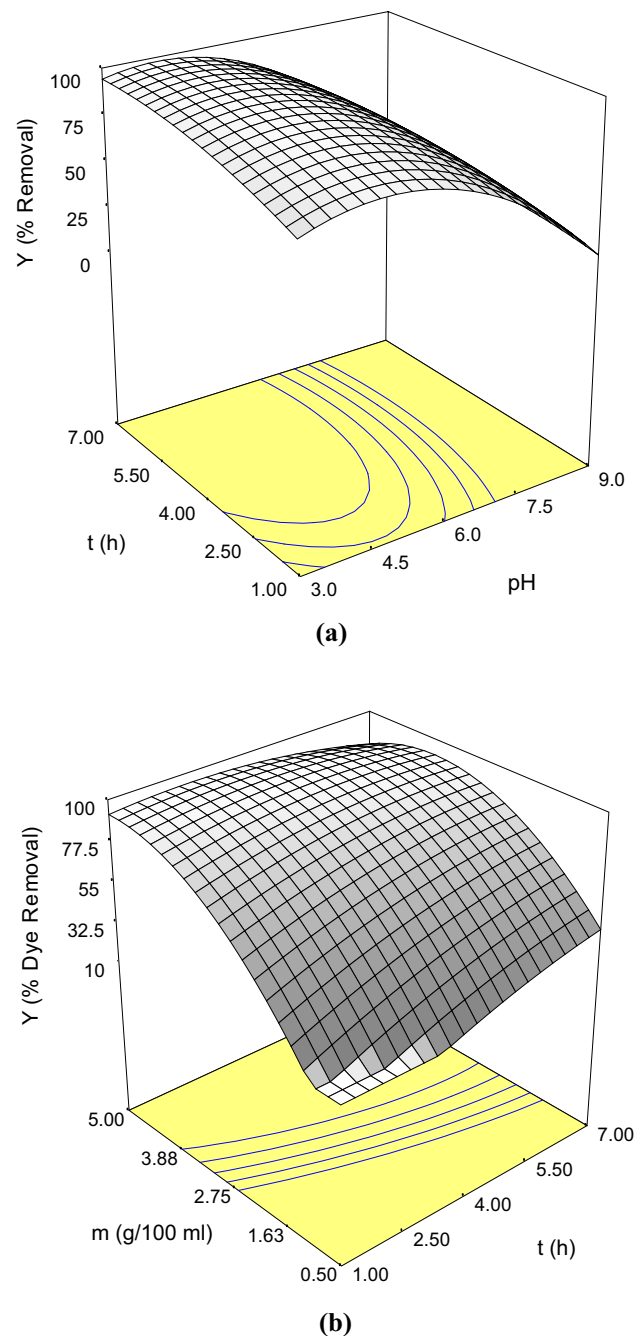
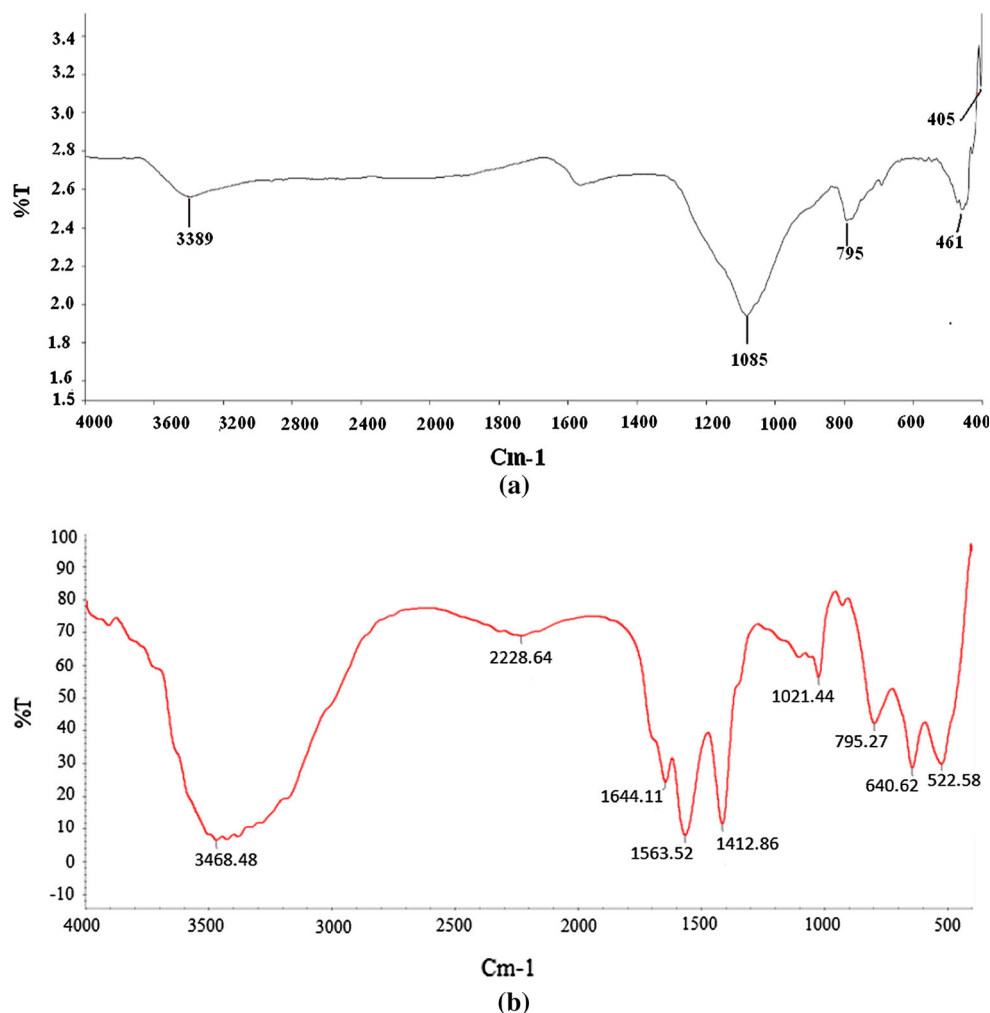


Fig. 1 Variation of CBSOL red LE wool dye % removal (Y) with **a** pH and adsorption time (t) and **b** adsorption time (t) and adsorbent dosage (m)

values, while at higher t values ($t > 5$ h), desorption of adsorbed dye onto CAC resulted in decreased response Y .

Due to the increase in m up to $m \approx 2.9$ g, availability of more adsorption sites increases surface for the adsorption



Fig. 2 FTIR spectra of **a** bare CAC **b** loaded CAC with dye

of dye on CAC, and hence, sharp increase in response Y is observed. Although for all $m > 2.9$ g active sites are increasing, but the adsorption is limited by the concentration of CBSOL LE red wool dye. Also, at all values of $m > 2.9$ g, increase in t values up to 5 h improves adsorption of CBSOL red LE wool but marginally, and for $t > 5$ h, desorption of CBSOL red LE wool dye was observed.

Optimization and adsorbent analysis

Adsorptive treatment of CBSOL red LE wool dye was optimized for maximizing response Y (% removal). The optimized values of operational parameters under the range of study were found to be $m = 2.92$ g/100 ml, $t = 6.75$ h and $\text{pH} = 3.95$. At these optimized parameters, response

Y was suggested to be 86 %. In order to validate the result suggested by optimization, three replicate runs were conducted with the optimized parameters and the average value of response Y was found to be 82.74 %.

However, experiments were conducted to validate the optimized process parameters and corresponding % color removal values generated for CBSOL red LE wool dye synthetic wastewater treatment with real textile industry wastewater. Real textile industry wastewater was collected from equalization tank of effluent treatment plant of a Woollen textile-based industry, located at Ludhiana, Punjab, India. For the purpose of validation, three experimental run were conducted at the optimized process parameters with the real textile wastewater. In order to calculate the % color removal, the original and treated samples were analyzed using a double beam UV–visible spectrophotometer



at $\lambda_{\max} = 473$ nm. The average % color removal was found to be 75.42 %, which is in good agreement with the CBSOL red LE wool dye synthetic wastewater % color removal.

At optimum condition, FTIR, SEM and EDX analysis of bare CAC and dye loaded CAC was conducted. The FTIR spectra of bare and dye loaded CAC are shown in Fig. 2. FTIR spectra of blank CAC were presented somewhere else. Sharp peaks at $\sim 1,085$, ~ 795 and ~ 461 cm^{-1} can be seen in FTIR spectra of blank CAC (Fig. 2a). Peak at $\sim 3,389$ cm^{-1} is due to O–H stretching mode of hydroxyl groups (Al-qodah and Shawabkah 2009). Peak at $\sim 1,085$ cm^{-1} may be accredited to existence of esters, carboxylic acids, alcohols and ethers, while the presence of the phenyl ring has their vibrations at ~ 795 cm^{-1} . Peak at ~ 461 cm^{-1} may be attributed to C–C stretching. Three peaks at $\sim 3,389$, $\sim 1,085$ and ~ 795 cm^{-1} can be seen to be affected after dye adsorption on CAC (Fig. 2b). This indicates that the functional groups at these wave numbers participate in the dye adsorption. Also, some new additional peaks at $\sim 2,228$, $\sim 1,644$, $\sim 1,563$, $\sim 1,412$, ~ 640 and ~ 522 can be observed. Peaks at ~ 522 – 640 are due to aromatic C–H vibration (Barkauskas and Dervinyte 2004). Bands around 1,600, 1,510, 1,460 and 1,426 cm^{-1} are assigned to aromatic skeleton vibrations (Carrott Suhas and Carrott Ribeiro 2007). However, peak at $\sim 1,644$ and $\sim 1,563$ is due to amides and nitro group stretching, respectively.

SEM (scanning electron microscope) and EDX (energy dispersive X-ray) analysis were conducted (JEOL JSM 6390 LV) in order to elucidate morphology and elemental distribution, respectively, of the bare CAC and loaded CAC (Fig. 3). Bare CAC shows porous structure having pores of different sizes with uneven surface, while for loaded CAC, the surface became smoother and non-porous due to the adsorption of dye (Fig. 3). EDX analysis revealed the presence of 36.08 % carbon, 38.22 % oxygen, 14.88 % silica and 6.83 % aluminum by weight percent in bare CAC. Dye loaded CAC showed 45.55 % carbon, 38.27 % oxygen, 9.35 % silica and 4.79 % aluminum. Except these elements, magnesium, sulfur, potassium, iron and calcium were found in trace amount in both bare and loaded CAC. It can be pointed out that loaded CAC showed higher carbon in comparison with bared CAC due to adsorbed dye.

Adsorption kinetics and diffusivity

Kinetics of adsorption provides useful data for the efficiency and practicability of the adsorption for large-scale

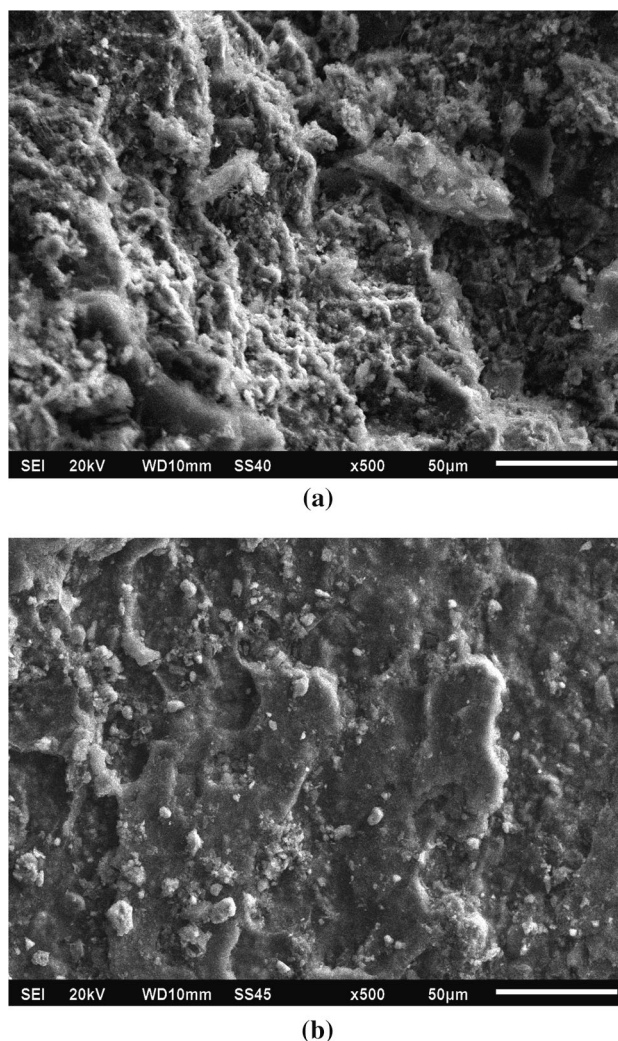


Fig. 3 SEM of **a** bare CAC **b** loaded CAC

operations. At optimized parameters ($m = 2.92$ g/100 ml, $t = 6.75$ h and $\text{pH} = 3.95$), kinetics of adsorption of CBSOL red LE wool dye from aqueous solution on CAC was studied using Pseudo-first-order and pseudo-second-order kinetic models. For this, solution of CBSOL red LE wool dye having initial concentration (C_0) 50 and 100 mg/l was kept in contact with the CAC at optimized parameters.

Pseudo-first-order model is given as (Malik 2003):

$$q_t = q_e [1 - \exp(-k_f t)] \quad (5)$$

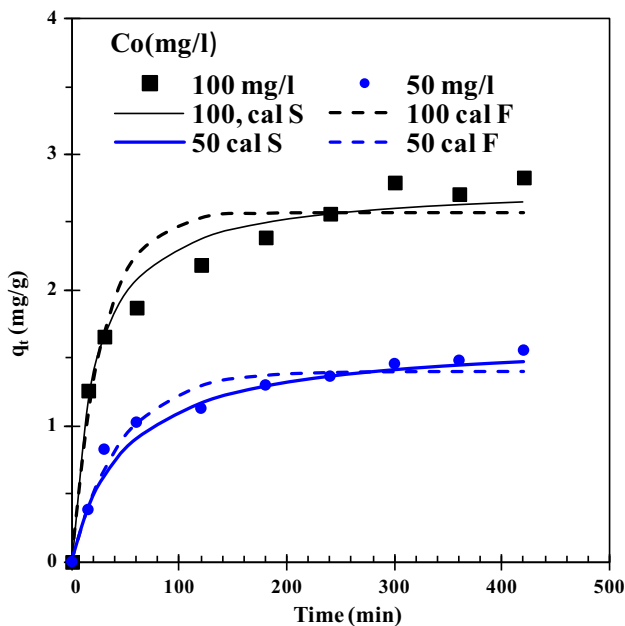
where q_t is amount of dye adsorbed (mg/g) at time (t) (min) and k_f is the rate constant of pseudo-first-order adsorption (min^{-1}).

Pseudo-second-order model is represented as (Blanchard et al. 1984; Ho and McKay 1999):



Table 3 Kinetic parameters for CBSOL red LE wool dye adsorption on CAC at optimized parameters ($m = 2.92$ g/100 ml, $t = 6.75$ h and pH = 3.95)

C_i (mg/l)	$q_{e, \text{exp}}$ (mg/g)	$q_{e, \text{cal}}$ (mg/g)	k_f (min^{-1})	R^2 (nonlinear)	MPSD
Pseudo-first-order model					
100	2.832	2.565	0.035	0.97	31.34
50	1.558	1.397	0.021	0.98	23.87
C_i (mg/l)	$q_{e, \text{exp}}$ (mg/g)	k_s (g/mg min)	h (mg/g min)	R^2 (nonlinear)	MPSD
Pseudo-second-order model					
100	2.780	0.018	0.137	0.99	17.60
50	1.647	0.012	0.033	0.99	19.51

**Fig. 4** Kinetics of CBSOL red LE wool dye adsorption on CAC at optimized parameters ($m = 2.92$ g/100 ml, $t = 6.75$ h and pH = 3.95). (Experimental results are shown by data points, solid line shows pseudo-second-order kinetic model fitting, and dotted line shows pseudo-first-order kinetic model)

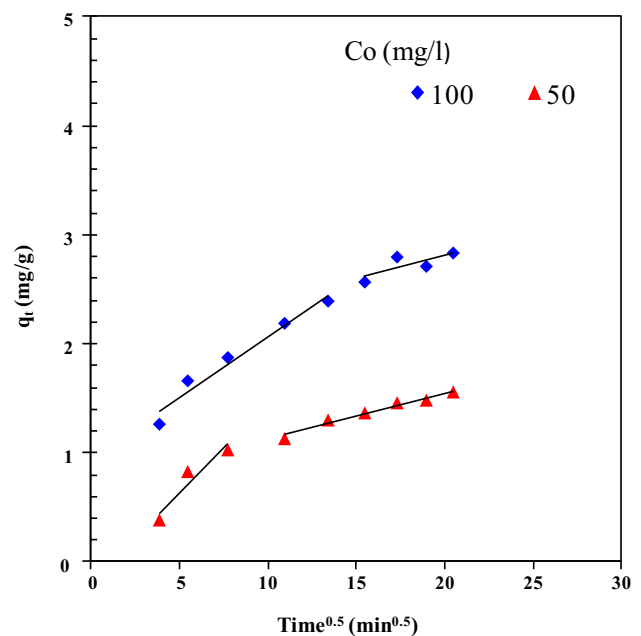
$$q_t = \frac{tk_s q_e^2}{1 + tk_s q_e} \quad (6)$$

where k_s is the rate constant of pseudo-second-order adsorption (g/mg min).

The initial adsorption rate, h (mg/g min), is defined as:

$$h = k_s q_e^2 \quad (7)$$

Nonlinear regression was followed to find the suitable model using Marquardt's percent standard deviation (MPSD) error function. The best-fit values of the model

**Fig. 5** CBSOL LE red wool dye adsorbed amount on CAC against square root of adsorption time (q_t vs. $t^{0.5}$)

parameters, coefficient of determination (R^2) and MPSD values are given in Table 3, and fitting is shown in Fig. 4.

As can be seen in Table 3, for pseudo-second-order model, R^2 values are higher than pseudo-first-order model fitting at both $C_0 = 50$ and 100 mg/l. Also, MPSD errors are smaller for pseudo-second-order model fitting within studied concentration range of (50–100 mg/l) CBSOL red LE wool dye. Therefore, it may be concluded that pseudo-second-order kinetic model best fits the adsorption kinetics.

From Table 3, it is observed that the q_e and k_s values were found to increase with an increase in C_0 . Therefore,



Table 4 Isotherm and thermodynamics parameters for CBSOL red LE wool dye adsorption on CAC at optimized parameters ($m = 2.92$ g/100 ml, $t = 6.75$ h and pH = 3.95)

Isotherms					
Freundlich					
		$q_e = K_F C_e^{1/n}$			
T (K)	K_F [(mg/g)/(l/mg) ^{1/n}]	n	R^2	SSE	
293	0.745	3.26	0.980	0.111	
303	0.898	2.87	0.998	0.014	
313	1.272	3.61	0.991	0.102	
Langmuir					
		$q_e = \frac{q_m K_L C_e}{1 + K_L C_e}$			
T (K)	K_L (l/mg)	q_m (mg/g)	R^2	SSE	
293	0.31295	2.255	0.985	0.114	
303	0.36032	2.755	0.974	0.337	
313	0.85875	2.822	0.988	0.273	
Redlich–Peterson					
		$q_e = \frac{K_R C_e}{1 + a_R C_e^\beta}$			
T (K)	K_R (l/g)	a_R (l/mg) ^{1/β}	β	R^2	SSE
293	1.865	1.6840	0.803	0.990	0.053
303	2.496	1.7137	0.812	0.990	0.128
313	7.777	4.8582	0.804	0.996	0.055
Thermodynamics parameters					
T (K)	$K \times 10^{-3}$ (l/kg)	ΔG_0 (kJ/mol)	ΔH_0 (kJ/mol)	ΔS_0 (J/molK)	
293	209.2775	−13.018	49.575	213.329	
303	366.9444	−14.877			
313	770.0478	−17.297			

the adsorption is limited by the C_0 of CBSOL red LE wool dye.

The initial rate of adsorption (h mg/g min) also increases with the increase in C_0 . This also supports inference that the adsorption is limited by the C_0 of CBSOL red LE wool dye in aqueous solution.

For the calculation of diffusivity, the diffusion equation based on Fick's second law can be approximated as (Pandey et al. 2013; Ruthven 1984).

$$q_t = 6q_e \left(\frac{Dt}{\pi} \right)^{0.5} \quad (8)$$

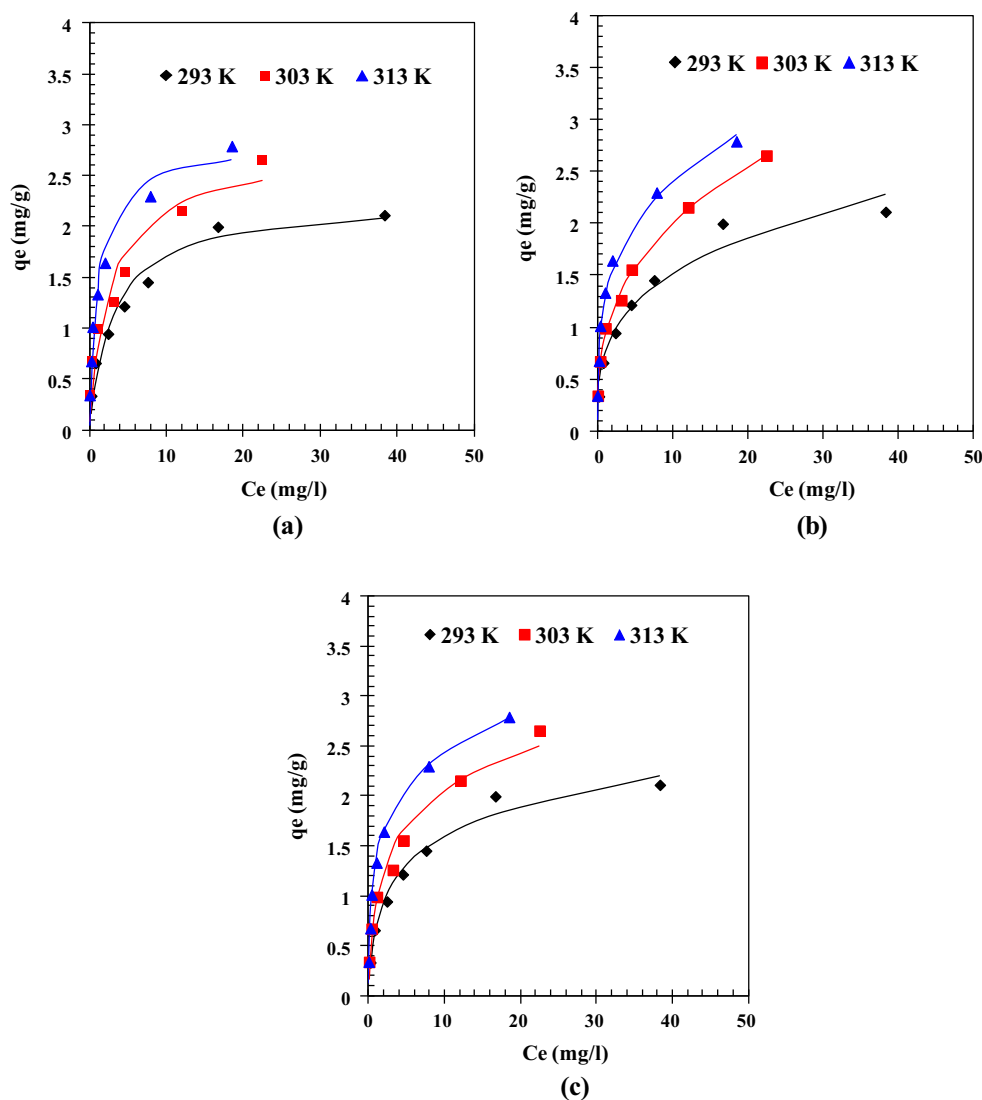
where D is diffusivity and q_t and q_e are the adsorption capacity at any time t and equilibrium, respectively.

For the calculation of diffusivity, identification of rate-controlling step in the adsorption process is necessary. Generally, mechanism of adsorption can be explored by plotting q_t versus $t^{0.5}$ (Pandey et al. 2013; Ruthven 1984;

Kushwaha et al. 2010; Rameshraj et al. 2012) (Fig. 5). This plot shows multi-linearity due to the various steps (film or external diffusion, pore diffusion, surface diffusion and adsorption on the pore surface) involved in the adsorption process. Figure 5 shows two linear portions over time range of study, which indicate that the rate of adsorption is controlled by more than one process (Kushwaha et al. 2010; Rameshraj et al. 2012; Srivastava et al. 2006). Since, in this study, good mixing condition was provided, film or external diffusion is very fast (line not shown in Fig. 5). The second linear portion shows gradual adsorption stage where the rate is limited by intra-particle diffusion (Ramaraju et al. 2013; Srivastava et al. 2006). The third portion shows that the adsorbate concentration in the solution becomes low, and the diffusion rate is decreased (Pandey et al. 2013; Crank 1965). Therefore, in this study, pore diffusion is the rate-limiting step and slope of second linear portion provides the diffusivity. The



Fig. 6 Equilibrium adsorption isotherms for CBSOL red LE wool dye adsorption on CAC at optimized parameters ($m = 2.92$ g/100 ml, $t = 6.75$ h and $\text{pH} = 3.95$) **a** Langmuir isotherm model **b** Freundlich isotherm model **c** R–P isotherm model. Experimental data points given by *symbols* and the *lines* predicated by isotherm model



calculated diffusivity with help of Eq. 8 was found to be 9.676×10^{-8} and 1.396×10^{-8} m^2/s , respectively, for C_0 of 50 and 100 mg/l, respectively.

Isotherm modeling and thermodynamics

Equilibrium isotherms supply the basic information for the design of adsorption systems. To study the adsorption isotherms for the CBSOL red LE wool dye adsorption onto CAC, three isotherms models Freundlich (Freundlich 1906), Langmuir (Langmuir 1918) and Redlich and Peterson (R–P) (Redlich and Peterson 1959) isotherms were analyzed.

To develop the equilibrium isotherms for CBSOL red LE wool dye adsorption onto CAC, experiments were conducted at different temperature (T) ranging from 293 to 313 K with $C_0 = 10$ –100 mg/l, at optimized parameters ($m = 2.92$ g/100 ml, $t = 6.75$ h and $\text{pH} = 3.95$). The data obtained were then fitted to various isotherm models using nonlinear regression with sum of squares error (SSE) function. The best-fit values of the model parameters, R^2 , and SSE values are given in Table 4, and fitting is shown in Fig. 6.

It was found that CBSOL red LE wool dye adsorption capacity increased with increase in T , concluding endothermic nature of adsorption (Fig. 6). This may be due to creation of new active surface on CAC surface at higher



T (Singh and Kushwaha 2014; Rameshraj et al. 2012). Also, the degree of diffusion within CAC structure is increased at higher T , enhancing the rate of intra-particle diffusion (Kushwaha et al. 2010; Aksu and Kabasakal 2004).

Table 4 shows the best-fit values of the model parameters, R^2 and SSE values. The Freundlich isotherm can be applied to heterogeneous surface. K_F and $1/n$ are the Freundlich constants and indicate adsorption capacity and intensity, respectively (Özer et al. 2009; Kushwaha et al. 2010). Increase in K_F with increase in T supports the endothermic nature of that CBSOL red LE wool dye adsorption onto CAC. The value of n greater than unity concludes the favorable adsorption conditions (Özer et al. 2009). Langmuir isotherm model suggests surface homogeneity of the adsorbent (Acar and Eren 2006). q_m and K_L of Langmuir isotherm model show affinity of adsorbent to adsorbate. The value of q_m and K_L was found to increase with increase in T , indicating strong bonding of CBSOL red LE wool dye molecule to CAC. In R–P isotherm, there are three parameters and is applicable to both the homogenous and heterogeneous surface. Values of β lie between 0 and 1 and signifying favorable adsorption. If $\beta = 1$, the physisorption is the main mechanism of adsorption, and $\beta = 1 - (1/n)$ shows adsorption is more heterogeneous in nature (Lakshmi et al. 2009). Table 4 shows β values are more nearer to $[1 - (1/n)]$, indicating adsorption of CBSOL red LE wool dye on CAC is heterogeneous in nature. R^2 values for Freundlich and R–P isotherms are higher (≈ 0.99) than Langmuir isotherm, and also error function SSE values are least for the Freundlich and R–P isotherms (Table 4). Therefore, Freundlich and R–P isotherms may be used to represent the equilibrium adsorption data.

Free energy change (ΔG_0), enthalpy change (ΔH_0) and entropy change (ΔS_0) were also estimated with the help of equilibrium adsorption constant (K) (Kushwaha et al. 2010; Acar and Eren 2006) using following equations:

$$\Delta G_0 = -RT \ln K \quad (9)$$

$$\ln K = \frac{\Delta S_0}{RT} - \frac{\Delta H_0}{R} \frac{1}{T} \quad (10)$$

where T is the absolute temperature (K), R is the universal gas constant ($8.314 \times 10^{-3} \text{ kJ mol}^{-1} \text{ K}^{-1}$) and $K = (q_e/C_e)$.

Table 4 shows thermodynamic parameters at various temperatures. Positive value of ΔH_0 and ΔS_0 shows the endothermic nature and favorable dye adsorption, respectively (Sharma 2011; Onal et al. 2006), and the negative value of ΔG_0 concludes spontaneous nature of adsorption

(Kushwaha et al. 2010; Srivastava et al. 2006; Lakshmi et al. 2009).

Conclusion

Optimization of adsorption process parameters for the removal of CBSOL LE red wool dye from aqueous solution by CCD under RSM was successfully applied, and quadratic model was found to best fit the experimental data. The values of R^2 and adjusted R^2 were found to be 0.894 and 0.80, respectively.

Kinetic study showed pseudo-second-order kinetic model best fitted the data, and adsorption is limited by the CBSOL red LE wool dye concentration.

Freundlich and R–P isotherms were found to represent the equilibrium adsorption data showing R^2 value 0.99. Increase in K_F value (Freundlich isotherm constant) with increase in T concluded endothermic nature of CBSOL red LE wool dye adsorption onto CAC. In R–P isotherm, the β values are nearer to $[1 - (1/n)]$, which indicates that the adsorption of CBSOL red LE wool dye on CAC is heterogeneous in nature. Positive value of ΔH_0 and ΔS_0 shows the endothermic nature and favorable dye adsorption, respectively, and the negative value of ΔG_0 concludes spontaneity of adsorption.

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References

- Acar FN, Eren Z (2006) Removal of Cu(II) ions by activated poplar sawdust (Samsun Clone) from aqueous solutions. *J Hazard Mater B* 137:909–914
- Aksu Z, Kabasakal E (2004) Batch adsorption of 2,4-dichlorophenoxy-acetic acid (2,4-D) from aqueous solution by granular activated carbon. *Sep Purif Technol* 35:223–240
- Al-qodah Z, Shawabkah R (2009) Production and characterization of granular activated carbon from activated sludge. *Braz J Chem Eng* 26(1):127–136
- Bansal S, Kushwaha JP, Sangal VK (2013) Electrochemical treatment of reactive black 5 textile wastewater: optimization, kinetics, and disposal study. *Water Environ Res* 85(12):2294–2306
- Barkauskas J, Dervinyte M (2004) An investigation of the functional groups on the surface of activated carbons. *J Serb Chem Soc* 69(5):363–375
- Behbahani M, Alavi Moghaddam MR, Arami M (2011) Technoeconomical evaluation of fluoride removal by electrocoagulation process: optimization through response surface methodology. *Desalination* 271:209–218



- Blanchard G, Maunaye M, Martin G (1984) Removal of heavy metals from water by means of natural zeolites. *Water Res* 18:1501–1507
- Carrott Suhas PJM, Carrott Ribeiro MML (2007) Lignin—from natural adsorbent to activated carbon: a review. *Bioresour Technol* 98:2301–2312
- Chakraborty S, Shamik C, Saha PD (2012) Batch removal of crystal violet from aqueous solution by H₂SO₄ modified sugarcane bagasse: equilibrium, kinetic, and thermodynamic profile. *Sep Sci Technol* 47:1898–1905
- Chiou M, Chuang G (2006) Competitive adsorption of dye metanil yellow and RB15 in acid solutions on chemically cross-linked chitosan beads. *Chemosphere* 62:731–740
- Chowdhury S, Saha P (2010) Sea shell powder as a new adsorbent to remove Basic Green 4 (Malachite Green) from aqueous solutions: equilibrium, kinetics and thermodynamic studies. *Chem Eng J* 164:168–177
- Crank J (1965) *The mathematics of diffusion*, 1st edn. Oxford Clarendon Press, London
- Freundlich HMF (1906) Over the adsorption in solution. *J Phys Chem* 57:385–471
- Gaikwad RW, Kanawade SM, Misal SA (2010) Low cost sugarcane bagasse ash as an adsorbent for dye removal from dye effluent. *Int J Chem Eng Appl* 1(4):310–318
- Geethakarthis A, Phanikumar BR (2011) Adsorption of reactive dyes from aqueous solutions by tannery sludge developed activated carbon: kinetic and equilibrium studies. *Int J Environ Sci Technol* 8(3):561–570
- Georgiou D, Melidis P, Aivasidis A (2002) Use of a microbial sensor: inhibition effect of azo reactive dyes on activated sludge. *Bioprocess Biosyst Eng* 25:79–83
- Gregory P (1986) Azo dyes: structure–carcinogenicity relationships. *Dyes Pigm* 7(1):45–56
- Güven G, Perendeci A, Tanyolac A (2008) Electrochemical treatment of deproteinated whey wastewater and optimization of treatment conditions with response surface methodology. *J Hazard Mater* 157:69–78
- Ho YS, McKay G (1999) Pseudo-second order model for adsorption processes. *Process Biochem* 34:451–465
- Izares P, Fabiolamartinez Carlosjimeanez, Justolobato Rodrigo M (2006) Coagulation and electro coagulation of wastes polluted with dyes. *Environ Sci Technol* 40:6418–6424
- Kushwaha JP, Srivastava VC, Mall ID (2010) Treatment of dairy wastewater by commercial activated carbon and bagasse fly ash: parametric, kinetic and equilibrium modelling, disposal studies. *Bioresour Technol* 101(10):3474–3483
- Lakshmi UR, Srivastava VC, Mall ID, Lataye DH (2009) Rice husk ash as an effective adsorbent: evaluation of adsorptive characteristics for Indigo Carmine dye. *J Environ Manag* 90:710–720
- Langmuir I (1918) The adsorption of gases on plane surfaces of glass, mica and platinum. *J Am Chem Soc* 40:1361–1403
- Lemlikchi W, Khaldi S, Mecherri MO, Lounici H, Drouiche N (2012) Degradation of Disperse Red 167 Azo Dye by bipolar electro-coagulation. *Sep Sci Technol* 47:1682–1688
- Malik PK (2003) Use of activated carbons prepared from sawdust and rice-husk for adsorption of acid dyes: a case study of Acid Yellow 36. *Dyes Pigm* 56:239–249
- Malik R, Ramteke DS, Wate SR (2006) Physico-Chemical and surface characterization of adsorbent prepared from groundnut shell by ZnCl₂ activation and its ability to absorb colour. *Indian J Chem Technol* 13:318–329
- Manoochehri M, Rattan VK, Khorsand A, Panahi H (2010) Capacity of activated carbon derived from agricultural waste in the removal of reactive dyes from aqueous solutions. *Carbon Lett* 11(3):169–175
- Montgomery DC (1996) *Design and analysis of experiments*, 4th edn. John Wiley & Sons, USA
- Myers RH, Montgomery DC (2002) *Response surface methodology: process and product optimization using designed experiments*, 2nd edn. John Wiley & Sons, USA
- Onal Y, Akmil-Basar C, Eren D, Saric-Ozdemir C, Depci T (2006) Adsorption kinetics of malachite green onto activated carbon prepared from tunçbilek lignite. *J Hazard Mater* B128:150–157
- Özer A, Gürbüz G, Calimli A, Körbahti BK (2009) Biosorption of copper(II) ions on *Enteromorpha prolifera*: application of response surface methodology (RSM). *Chem Eng J* 146:377–387
- Pandey NK, Velavendan P, KamachiMudali U, Natarajan R (2013) Adsorption of di-butyl phosphate on activated alumina: equilibrium and kinetics. *Desalin Water Treat*. doi:10.1080/19443994.2013.839397
- Parvathi C, Maruthavanan T (2010) Adsorptive removal of Megenta MB cold brand reactive dye by modified activated carbons derived from agricultural waste. *Indian J Sci Technol* 3(4):408–410
- Rafatullaha M, Sulaimana O, Hashima R, Ahmad A (2010) Adsorption of methylene blue on low-cost adsorbents: a review. *J Hazard Mater* 177:70–80
- Ramaraju B, Kumar Reddy PM, Subrahmanyam C (2013) Low cost adsorbents from agricultural waste for removal of dyes. *Environ Prog Sustain Energy* 33:38–46
- Rameshraj D, Srivastava VC, Kushwaha JP, Mall ID (2012) Quinoline adsorption onto granular activated carbon and bagasse fly ash. *Chem Eng J* 181–182:343–351
- Redlich O, Peterson DL (1959) A useful adsorption isotherm. *J Phys Chem* 63:1024–1026
- Ruthven DM (1984) *Principles of adsorption and adsorption processes*. Wiley, New York
- Seey TL, Noordin MJ, Kassim M (2012) Characterization of mangrove bark as a potentially low-cost adsorbent for reactive dye removal from aqueous solutions: equilibrium, mechanisms and kinetics. *Int J Pure Appl Sci Technol* 9(1):9–19
- Sharma YC (2011) Adsorption characteristics of a low cost activated carbon for the reclamation of colored effluents containing malachite green. *J Chem Eng Data* 56:478–484
- Singh S, Kushwaha JP (2014) Tannic acid adsorption/desorption study onto/from commercial activated carbon. *Desalin Water Treat* 52:3301–3311
- Soo EL, Salleh AB, Basri M, Rahman RNZA, Kamaruddin K (2004) Response surface methodological study on lipase-catalyzed synthesis of amino acid surfactants. *Process Biochem* 39:1511–1518
- Srivastava VC, Swamy MM, Mall ID, Prasad B, Mishra IM (2006) Adsorptive removal of phenol by bagasse fly ash and activated carbon: equilibrium, kinetics and thermodynamics. *Colloid Surf A Physicochem Eng Asp* 272:89–104
- Suteu D, Malutan T, Bilba D (2011a) Agricultural waste corn cob as a sorbent for removing reactive dye orange 16: equilibrium and kinetic study. *Cellulose Chem Technol* 45(5–6):413–420
- Suteu D, Zaharia C, Malutan T (2011b) Removal of Orange 16 reactive dye from aqueous solutions by waste sunflower seed shells. *J Serb Chem Soc* 76(4):607–624



Weng CH, Lin YT, Tzeng TW (2009) Removal of methylene blue from aqueous solution by adsorption onto pineapple leaf powder. *J Hazard Mater* 170:417–424

Yazdani M, Mahmoodi NM, Arami M, Bahrami H (2012) Isotherm, kinetic, and thermodynamic of cationic dye removal from binary system by Feldspar. *Sep Sci Technol* 47(2–12):1660–1672

